

Experiment to Test the pH of Different Solutions

Page 1

Name		School	
Class		Date	

Results

Record your results a table:

Solution	pH

Conclusion

Which of the solutions gave an acid pH?	
Is this the result you expected?	
Which of the solutions was alkaline?	

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Page 2

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How can you relate this result to the use of the liquid?	
Which of the solutions were neutral? Did you expect these results? Explain.	
Further Investigation To answer the following questions you can either try out the procedure in the laboratory or, using your knowledge and the information given, make hypotheses about what might happen.	
Neutralization If a drop of alkali (pH12) is added to an acid (pH2) in a test tube you will find that, after mixing the contents of the test tube, the pH value of the acid has changed. You might read pH3. If you continue to add the alkali drop by drop, mixing and testing after every added drop, you will observe that the pH moves closer to pH7. If you are lucky you will observe a pH of 7. This is neutralization. If you are less fortunate you will find that the pH moves above pH7.	
Why should the pH move above pH7?	
What would you now do to reach pH7?	
Dilution If you take 1cm ³ of a pH2 acid, add 9cm ³ of distilled water and mix them well, what would you expect to happen to the pH of the diluted acid? Explain your answer.	