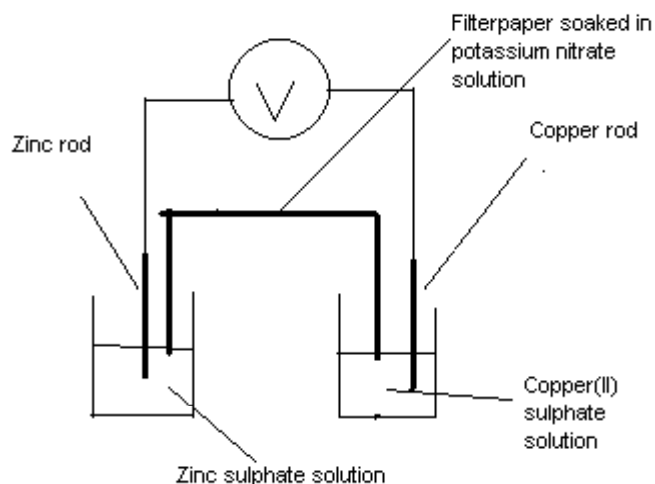


## Electricity from chemicals

Name:	Class:
Teacher:	Date:

Set up the apparatus below:



1. Observe the voltmeter. Explain what happens and why.

2. Repeat the experiment using iron and iron(II)sulphate solution instead of copper and copper(II) sulphate. Explain what happens and why?

3. For each of the two experiments:

a. Which metal in each case is more able to form ions?

b. Which metal will become the negative electrode of the cell and which the positive?

c. Draw a simple diagram of the zinc-iron cell showing the direction of the movement of electrons.

4. Repeat the experiment with other pairs of metals, state each time which one is to be the negative electrode.

Extra: Investigate how you can make simple cells using a lemon and potatoes and two different metals.