Reduction of Lead Oxide

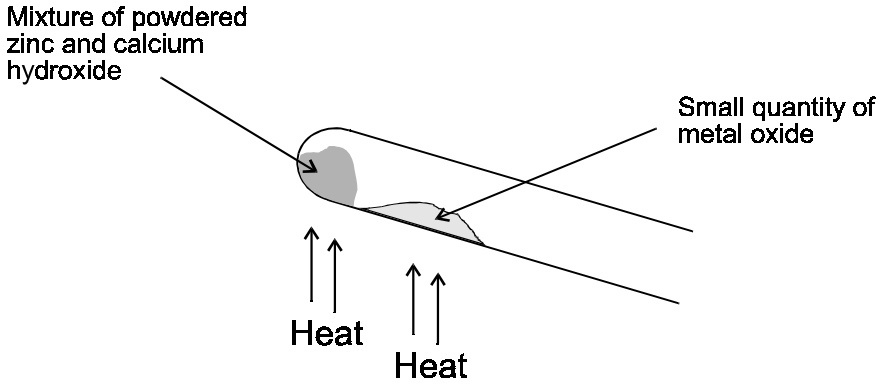
The oxygen can be removed from lead oxide by heating the oxide in the presence of a substance which is _____ than lead.

This process is called _____ the oxide.

For example carbon is more reactive than lead and so it can remove the oxygen from lead oxide, forming _____ and _____.
Hydrogen is more reactive than lead and so it can remove the oxygen from lead oxide, forming _____ and _____.

Experiment to reduce some lead oxide

The mixture of powdered zinc and calcium hydroxide produces _____ when heated.



Method

Heat the metal oxide, moving the test tube slowly all the time. Then move the test tube so that the mixture of zinc and calcium hydroxide is in the hottest part of the flame.

Repeat the above procedure and observe the oxide. During the heating, the test tube should be *angled down slightly*, as shown in the diagram, so that _____.

Observations

_____

_____

_____

_____

Conclusion

_____

_____

_____

_____

Describing the reaction by a chemical equation

1. Describe the reaction in words.
   Before the reaction we had lead oxide (_____) and hydrogen (___).

   The _____was removed from the oxide.

   After the reaction we had _____and _____ .

2. Basic equation

3. This equation *is/is not* balanced.

4. Complete equation